

Thermodynamic Incompatibility and Complex Formation in Pectin/Caseinate Mixtures

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The instability of polysaccharide/protein mixtures occurs because of either thermodynamic incompatibility or complexation. We studied which instability mechanism dominated given the external conditions. Therefore the effect of temperature, pH, and biopolymer concentration on the phase separation of pectin/caseinate mixtures was investigated. At pH > 6, thermodynamic incompatibility with spinodal decomposition was observed in pectin/caseinate mixtures resulting in the formation of water-in-water emulsions in intermediate stages of the phase separation process. The demixing rate of these emulsions and appearance of two macroscopic phases (lower phase enriched with caseinate and upper phase with pectin) was retarded when the pectin concentration increased or when the storage temperature decreased due to a higher viscosity of the mixtures at those conditions. As the pH of the mixture was lowered below 6, pectin accumulated in the caseinate-rich phase. Complexation of pectin and caseinate led to the formation of microparticles ($\sim 3 \mu\text{m}$), whose shape depends on the biopolymer concentration ratio and rate of acidification. These pectin/caseinate particles do not coalesce and are insensitive to salt addition.

1. Introduction

Aqueous mixtures of polysaccharides and proteins are usually unstable above a certain biopolymer concentration and phase-separate either because of repulsive (segregative) or attractive (associative) interactions.^{1–4} In the first case, also called thermodynamic incompatibility, the two biopolymers do not carry opposite net charges, and after stirring is stopped, the dispersion separates into two phases, each enriched by one of the polymers. In the case of complexation, the two biopolymers attract each other because they carry opposite charges, and after demixing, the system consists of a phase containing biopolymer complexes and another phase consisting mainly of solvent. We demonstrate that both types of phase instabilities can be found in aqueous mixtures of pectins and caseinates.

Because of the wide applicability of polysaccharide/protein mixtures, the interest in identifying interactions between them has been growing to provide optimum food quality, in terms of texture and stability, and also innovative products in cosmetic and pharmaceutical industries.

Pectin is an anionic linear polysaccharide extracted from citrus peel or apple pomace. Its chemical structure is mainly characterized by D-galacturonic acid units linked together by α -(1–4) glycosidic linkages, occasionally interrupted by α -(1–2) L-rhamnopyranosyl residues. Caseinates are milk proteins whose monomers are complex amphiphilic copolymers with a strong tendency for micellar self-assembly. In milk, caseins are associated with calcium phosphate forming casein micelles (diameter $\approx 200 \text{ nm}$).⁵ After removing calcium phosphate, casein micelles dissociate into casein monomers and/or casein submicelles.⁶

Mixtures of pectin and casein(ate) in aqueous solution are known to be incompatible at pH 6.5 above a certain biopolymer

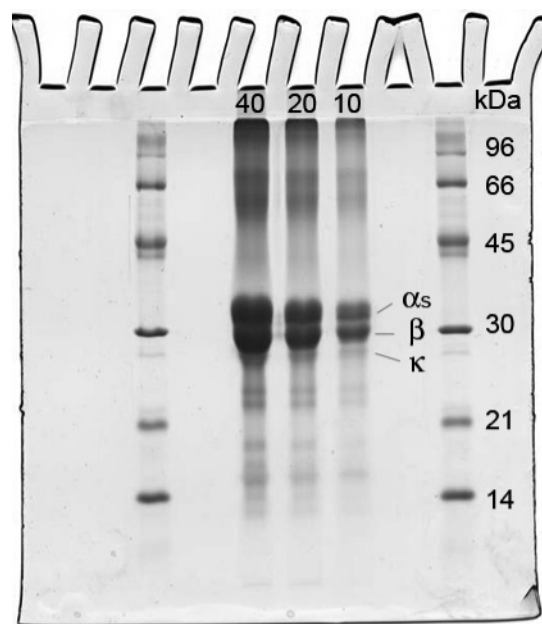


Figure 1. Image of the SDS-PAGE gel with separated sodium caseinate fractions.

concentration.⁷ At lower pH values, pectin has been used to stabilize acidified dairy beverages due to pectin adsorption onto casein micelles, preventing the aggregation of casein micelles during the acidification process. Much has been published about the interaction of pectin with casein micelles in acidified milk drinks,^{8–11} but very few studies were done on pectin/caseinate mixtures at low pH's in calcium-free aqueous solutions.^{12,13} None of them, however, used pectin concentrations above 1 wt %, because their interest was focused on milk drink stabilization and that requires just small amounts of pectin.

In this work we study the phase behavior of aqueous pectin plus caseinate mixtures in a wide pH and biopolymer concentra-

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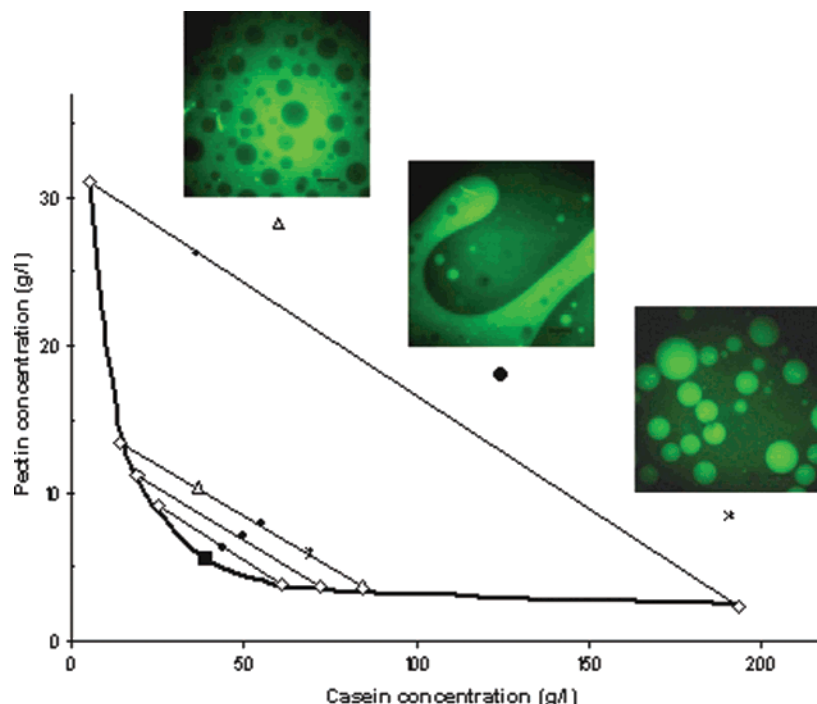


Figure 2. Phase diagram of pectin/casein mixtures at pH 7 in 0.01 M phosphate buffer: diamonds, binodal points; circles, initial concentrations; square, critical point; thick full curve, binodal; thin full line, tie-lines. Images show mixtures with different volumes of polymer phases for the same tie-line: upper, 75% pectin/25% caseinate; middle, 50% pectin/50% caseinate; lower, 25% pectin/75% caseinate. Bright areas correspond to the pectin-rich phase.

tion range. This enables optimizing the use of pectin/caseinate mixtures in applications (see below). Therefore, a phase diagram was obtained at high pH's, where thermodynamic incompatibility takes place, to determine the concentrations where stable and unstable mixtures can be found. Inside the unstable region, water-in-water emulsions were formed, and the stability of these emulsions was investigated as a function of biopolymer concentration and storage temperature. At low pH's, we investigated the complex formation between pectin and caseinate and the conditions to obtain large spherical particles of the complex. The process to obtain such particles, which is simple and performed in mild conditions, could be used as a microencapsulation technique for pharmaceutical compound encapsulation or for encapsulation of flavors and odors in food products. The advantage of working with pectin and caseinate is that they are biodegradable polymers with recognized alimentary use and no toxic effects in their utilization are expected.

2. Materials and Methods

2.1. Materials. High methoxy citrus pectin GENU type USP-B with a 68% degree of esterification was kindly donated by CP Kelco Limeira, Brazil. CP Kelco indicates that the molar mass is between 110 and 130 kg/mol. Commercial bovine milk sodium caseinate was a gift from Kauffmann & Co., Germany. The caseinate powder contains 92–94% protein and has a maximum humidity of 12%. Fluorescein-5-isothiocyanate (FITC) and rhodamine B were purchased from Sigma-Aldrich, and dibutyltin dilaurate and pyridine from Fluka. All of the other reagents used in this study were of analytical grade.

2.2. Methods. **2.2.1. Casein Characterization.** A semiquantitative characterization of casein was performed using the polyacrylamide gel electrophoresis (PAGE) technique, according to the method of Laemmli¹⁴ in a discontinuous gel system composed of 5% and 12.5% acrylamide stacking and running gels, respectively. Sodium caseinate solution at pH 7 was diluted in 0.025 M Tris-HCl buffer pH 6.8 containing 20% glycerol, 0.02% bromophenol blue, and 0.2 M

dithiothreitol (DTT) and boiled for 5 min at 100 °C. Samples with 40, 20, and 10 mg of protein were loaded into a vertical electrophoresis cell (Hoefer SE 600, Amersham Biosciences) at 12 mA and ran until the dye front reached the bottom of the gel. The electrode buffer was 0.025 M Tris-HCl pH 6.8, 0.192 M glycine, and 0.1% sodium dodecyl sulfate (SDS). The gel was stained with Coomassie brilliant blue G-250. First it was incubated for 1 h in a 40% (v/v) ethanol and 10% (v/v) acetic acid solution and then immersed for 12 h in 0.1% w/v Coomassie G-250, 2% w/v phosphoric acid, 10% w/v ammonium sulfate, and 20% v/v methanol. For destaining, 25% v/v methanol was used, and the images were acquired with ImageScanner II (Amersham Biosciences).

2.2.2. Labeling. To visualize pectin and caseinate with confocal scanning light microscopy (CLSM), they were labeled with fluorescent dyes. Pectin was labeled covalently with FITC according to the procedure described by Tromp et al.¹⁵ Pectin (1 g), FITC (0.14 g), pyridine (400 μ L), and dibutyltin dilaurate (80 μ L) were dissolved in 100 mL of dimethylsulfoxide (DMSO). The reaction mixture was heated for 4 h at 100 °C, poured into isopropyl alcohol (400 mL), and centrifuged for 55 min at 9500g at 20 °C. Subsequently, the polysaccharide sample was redissolved in 200 mL of hot water and dialyzed against 0.1 M NaCl solution (4 \times 4 L) and distilled water (2 \times 4 L). In this manner an excess of unbound dye and salt was removed. After that, the sample was freeze-dried.

To calculate the number of FITC molecules per galacturonic acid molecule, first a calibration curve for FITC was constructed. An amount of 10 mg of FITC was dissolved in 100 mL of DMSO. Dilutions of this solution in 0.01 M phosphate buffer were used to construct the calibration curve, measured at a wavelength of 490 nm, from which the molar extinction coefficient (ϵ) was calculated, giving $\epsilon = 73\,789\text{ L mol}^{-1}\text{ cm}^{-1}$. To determine the number of dyes per pectin monomer, we dissolved a known amount of pectin-FITC in buffer and detected the absorption of the solution at 490 nm. The dye concentration then follows from absorption plus the obtained extinction coefficient. We found that each milligram of labeled pectin contained 20.45 μ g of FITC. Considering that pectin mainly consists of galacturonic acid monomers with a molar mass close to 200 g/mol, the number of FITC molecules/galacturonic acid molecules was estimated to be about 0.01. The fact

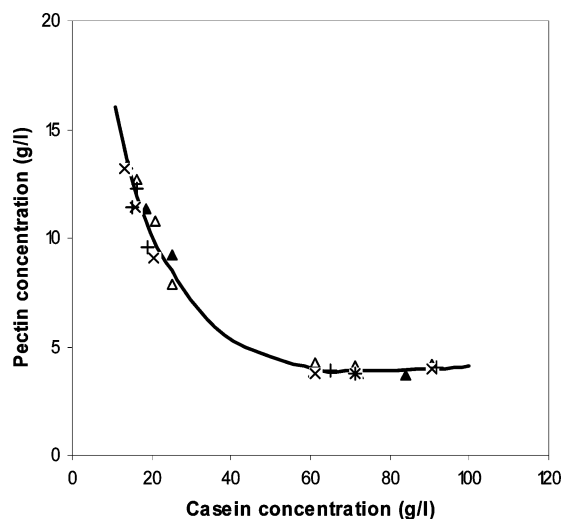


Figure 3. Effect of temperature on the phase diagram of the pectin/caseinate mixtures at pH 7. Samples were incubated at 45 °C (triangles, two series of experiments), 20 °C (x), and 5 °C (+).

that we have only one dye molecule per hundred sugar monomers and that the phase behavior does not change with the small amount of labeled pectin used implies that we cannot expect large effects of labeled pectin on the physical properties.

Caseinate was labeled noncovalently with rhodamine B by adding one drop of a 0.02% solution to 10 mL of biopolymer mixture.

2.2.3. Phase Diagram Determination. Pectin and sodium caseinate were dispersed separately in 0.01 M phosphate buffer pH 7.0 containing 0.03 wt % sodium azide (NaN_3) to prevent bacterial contamination. Stock solutions of 1.76 wt % pectin and 10 wt % caseinate were used to prepare mixtures as to establish the phase diagram. Equal volumes of the diluted solutions were mixed, and the pH was adjusted to 7.0 with 4 M NaOH. One gram of the mixtures was weighed in glass tubes and left overnight to phase-separate at 45 °C. The concentrations of the biopolymers in both phases were determined spectrophotometrically. Two series of samples were prepared, one containing labeled pectin (50 mg of pectin-FITC/g of unlabeled pectin) and another with only unlabeled pectin for caseinate quantification. No difference of the phase volumes was observed between the labeled and the unlabeled samples. To measure pectin and caseinate concentrations, calibration curves for the two polymers were constructed at pH 7.0. The absorbance values were taken at a wavelength of 490 nm for pectin-FITC ($\epsilon = 0.21 \text{ L g}^{-1} \text{ cm}^{-1}$). To quantify the amount of caseinate, the Bradford reagent was used,¹⁶ and the absorbance values were taken at 595 nm, where $\epsilon = 0.63 \text{ L g}^{-1} \text{ cm}^{-1}$.

The most upper tie-line in the phase diagram was obtained from solutions of 6 wt % pectin and 6 wt % caseinate. In this case, the phase separation took a long time, and therefore, 1 g of the mixtures were weighed in small vials and centrifuged in a Microfuge18 centrifuge (Beckman Coulter) at 2296g for 1 h. Also, as the lower phase in this sample was very viscous and difficult to pipet, the concentration in that phase was obtained by subtracting the amount of the biopolymers (pectin and caseinate) in the upper phase from the total amount. The volumes were determined by marking the interface position on the outside of the tubes and, after emptying, by weighing the amount of water corresponding to the marked volumes.

2.2.4. Effect of pH on Pectin/Caseinate Mixtures. Equal volumes of 6 wt % pectin and 6 wt % caseinate solutions were mixed, and the pH was adjusted to 8.0 with 4 M NaOH. Under magnetic stirring, the mixture was acidified with 1 M citric acid to various pH values between 8 and 3. At each pH, 5 g of the mixtures were collected and centrifuged in a TJ-25 Centrifuge (Beckman Coulter) at 5000g for 3 h. Again, two series of samples were prepared for quantification of pectin and caseinate in the upper phases (according to subsection 2.2.3). Calibration curves were established at every pH. The concentrations in the lower

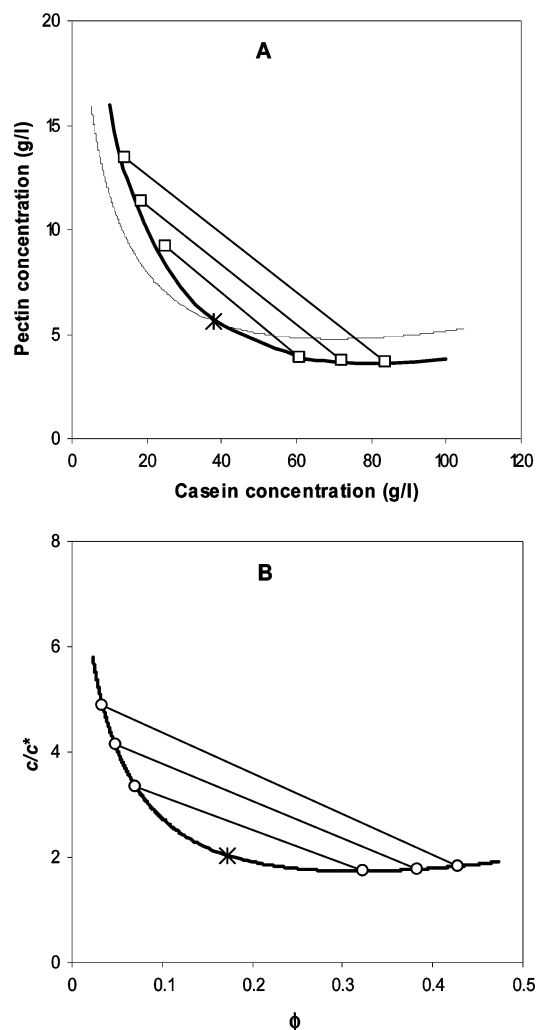


Figure 4. Comparison of experimental pectin/caseinate phase diagram (A) at pH 7 with theoretical phase diagram (B) for mixtures of hard spheres plus polymer chains with excluded volume interaction.²⁴ Symbols identify experimental (squares) and a few representative theoretical (circles) binodals connected through tie-lines. The theoretical prediction is also represented in panel A by the thin curve. Asterisks identify the critical points.

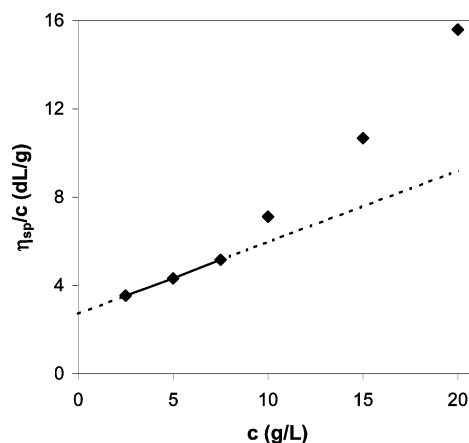


Figure 5. Specific viscosity η_{sp} over pectin concentration c as a function of pectin concentration measured at pH 7. The intercept yields the intrinsic viscosity $[\eta]$: measurements, data points; straight line, linear fit at low concentrations.

phase were obtained by subtracting the amount of the biopolymers in the upper phase from the total amount present in the mixture. The volumes were measured as described also in section 2.2.3.

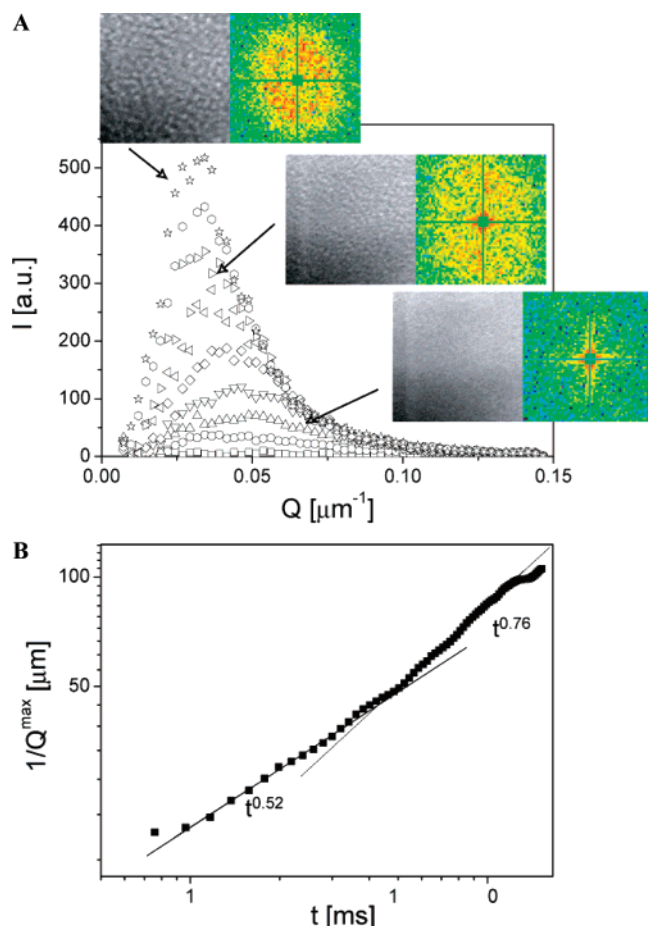


Figure 6. Spinodal decomposition for the 50% pectin/50% caseinate sample. (a) Intensity of the Fourier transform of the fluorescence microscopy images. The open squares represent the data set of intensities taken 0.3 s after mixing, and the stars were data taken 1.7 s after mixing. The insets show the real space (field of view is 400 μm) and Fourier transforms at the indicated times. (b) Structure size ($1/Q_{\text{max}}$) as a function of time.

2.2.5. Stability. The stability of pectin/caseinate mixtures at pH 7.0 with different polymer ratios was studied at 5, 15, 30, and 45 $^{\circ}\text{C}$. One gram of each mixture was weighed in glass tubes and closed to avoid solvent evaporation. The tubes were kept in the thermostatic bath or fridge for 1 week. The samples were checked every 24 h. When the meniscus between the two phases appeared, the sample was considered demixed.

2.2.6. Rheology. Rheological experiments of concentrated samples were performed using a rheometer Rheometric Scientific (Ares) in the controlled strain mode with Couette geometry (bob 32 mm and cup 34 mm). The measurements were made at 5, 15, 30, and 45 $^{\circ}\text{C}$. A solvent trap was used to avoid evaporation.

The intrinsic viscosity of pectin was determined at 25 $^{\circ}\text{C}$ by using an ordinary Ostwald viscosimeter. An amount of 15 mL of each pectin solution, prepared in a phosphate buffer with pH 7, was poured into the viscosimeter. The flow time of the samples was measured at different concentrations (0.25, 0.5, 0.75, 1, 1.5, and 2 wt %). The dynamic viscosity η follows from the flow time multiplied by the density. The density (ρ) was measured using a pycnometer.

2.2.7. Microscopic Observations. Transmission microscopy and CSLM observations were carried out on a Zeiss Axioplan 2 with a 40 \times oil immersion objective lens. The scanning head was a multipoint confocal scanner VTinfinity from VisiTech. The excitation lengths for rhodamine and FITC were 568 and 488 nm, respectively. Emission maxima were detected at wavelengths of 625 and 518 nm.

The microscopic observations of the phase separation were done on a Zeiss Axioplan 2 with a 10 \times objective lens. Samples containing

rhodamine were mixed in a vortex, and 50 μL was immediately pipetted on a glass slide. With a spatula, the samples were sheared continuously until the sample was well-mixed, and immediately after cessation of mixing, images were taken sequentially.

Image treatment was done with home-written IDL programs, using the standard IDL Fast-Fourier Transform routine.

3. Results and Discussion

3.1. Casein Characterization. The PAGE traces for caseinate solution, shown in Figure 1, are very similar to already reported bovine casein PAGE profiles.^{17–19} Two major bands appeared after protein separation, which correspond to α_s -casein (approximately 32 kDa) and β -casein (approximately 29 kDa). The amounts of these two fractions were similar. The third fraction, κ -casein, appears as a thin and smooth trace (approximately 27 kDa), indicating a smaller amount. This result agrees with the ratio 4:1:4:1.3 of the fractions $\alpha_s/\alpha_{s2}/\beta/\kappa$ generally found for bovine caseins.⁶

3.2. Phase Diagram. When pectin and caseinate are brought together in aqueous solution the mixture may phase-separate into two phases: a lower phase containing mainly protein and an upper phase enriched in polysaccharide. To determine the phase coexistence concentrations at which pectin/caseinate mixtures undergo phase separation we constructed a phase diagram, as shown in Figure 2, at pH 7. The solid curve corresponds to the binodal, which distinguishes the compatible (one-phase) region from the incompatible (two-phase) region in the phase diagram. Below the binodal, pectin/caseinate mixtures are stable. As the concentration of the biopolymers is raised above the binodal, the system starts to phase-separate. The compatible region is, therefore, small compared to the incompatible one, as expected for polysaccharide/protein mixtures²⁰ and according to previous work of Antonov et al.⁷

In the phase diagram the concentrations of the coexisting phases after phase separation are linked by tie-lines. As the mixture is diluted, while taking care that the volumes of the two coexisting phases remain approximately equal, the tie-lines become shorter until the coexisting phase compositions merge at the critical point. At relatively high pectin concentrations along a tie-line, the system is microscopically characterized by droplets of caseinate (dark domains) in a continuous pectin phase (light domains) in Figure 2. If we take the opposite direction toward higher caseinate concentrations, then there is a phase inversion, and pectin droplets become surrounded by a continuous phase enriched in caseinate. At intermediate positions along a tie-line, where the volume fractions of both phases are similar, a bicontinuous system is observed.

A mixture of caseinate and pectin may be considered as a colloid/polymer mixture in a common solvent. At the temperature studied, caseinates (mainly β - and α -caseins) form spherical micelles with a diameter of 30 nm.²¹ Pectin molecules in aqueous solution can be regarded as polymer chains. In case polymers (pectin) do not adsorb onto spherical particles (micellar caseinates), the non-adsorbing polymers are depleted from the regions surrounding the spheres. Such regions are called depletion zones.^{22,23} When the protein particles approach each other sufficiently close, the depletion zones overlap. Thus, the distance between the two spheres is smaller than the diameter of the polymer particles, and therefore no polymers can enter the gap between them. Such a lack of polymers leads to an osmotic pressure imbalance that causes an effective attraction between the spheres. If this attraction is sufficiently strong, then the system will phase-separate; the spheres and non-adsorbing

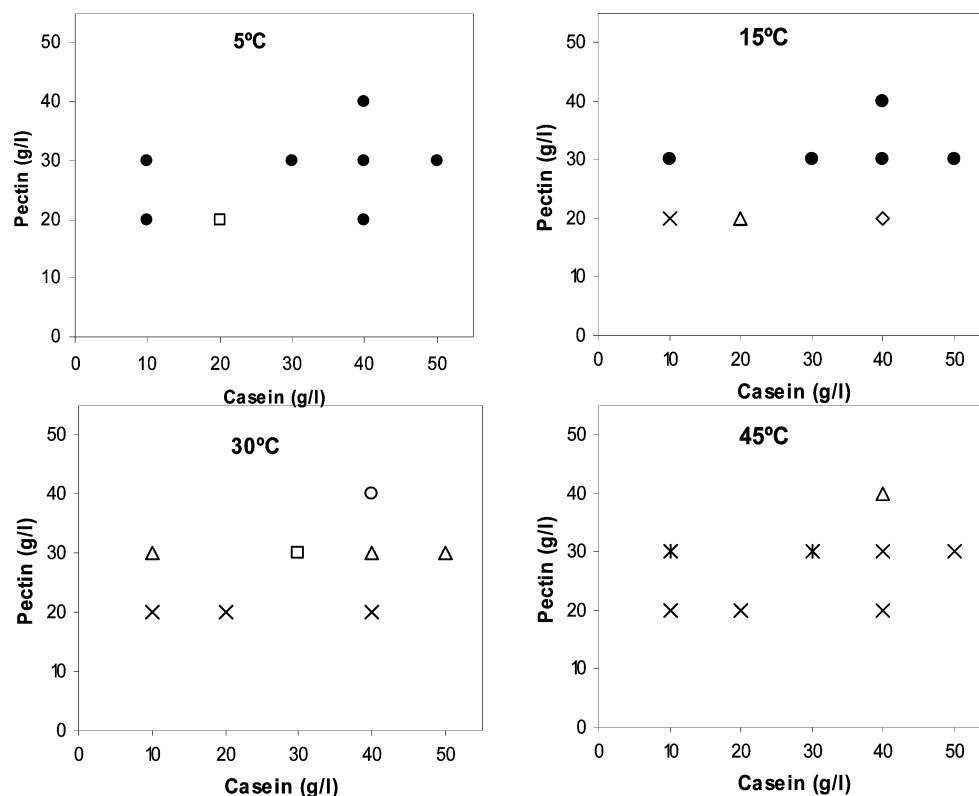


Figure 7. Time of phase separation at pH 7 at different temperatures: crosses, 24 h; asterisks, 48 h; triangles, 72 h; diamonds, 96 h; squares, 120 h; open circles, 144 h; filled circles, stable until 1 week.

polymers are thermodynamically incompatible. Thermodynamically, the driving force is purely entropic, and therefore, no temperature dependence of the phase coexistence curve is expected in this case. As demonstrated in Figure 3, mixtures incubated at three different temperatures (45, 20, and 5 °C) had similar results for the coexisting phase concentrations after phase separation.

In Figure 4, we compared the experimental coexistence curve for pectin/caseinate mixtures at pH 7 (panel A) to a theoretical prediction²⁴ for mixtures of hard spheres plus polymer chains with excluded volume interaction between the segments (panel B). The theoretical phase diagram is only a function of $q = R_g/a$ with hard sphere radius a and the radius of gyration of the polymer chains R_g . Caseinates are mainly present as spherical micelles with a radius $a \approx 15$ nm.²¹ For the pectin polymers, we have no information of the size of the chains. Therefore, we match the critical point of the theoretical prediction to the experimental critical point. The critical caseinate concentration is close to 40 g/L (Figure 2). The volume fraction ϕ of the spherical micelles can be converted to the caseinate concentration by using the voluminosity of caseinate micelles, being close to 4 mL/g.²¹ We therefore set $q = R_g/a = 1.5$ in the calculations because this corresponds to a critical volume fraction at $\phi = 0.17$. The polymer concentration at the critical point is found in terms of c/c^* , where c^* is the overlap concentration. Theoretically, we find $c/c^* \approx 2$ at the critical point for $q = 1.5$. We thus ensure $c/c^* = 2$ corresponds to the pectin concentration at the critical point determined experimentally, which is 5.6 g/L. This means c^* would be close to 2.7 g/L. Because $c^* \approx 1/[\eta]$, where $[\eta]$ is the intrinsic viscosity, the intrinsic viscosity would be on the order of 0.4 L/g.

We measured the viscosity of the pectins in the solution that we used in the dilute regime. In Figure 5 we plot the specific viscosity over the pectin concentration as a function of pectin

concentration. The intrinsic viscosity $[\eta]$ follows from the intercept and is 0.27 L/g, close to the result of Kar and Arslan.²⁵

We plot the theoretical phase coexistence curve in Figure 3 (panel B), including a few representative tie-lines. The resulting theoretical phase diagram is very similar to the experimental phase diagram, which supports the idea that the phase behavior is only driven by entropic forces. It is noted that a prediction for hard spheres plus ideal (non-interacting) chains looks differently; the binodal then looks more like a $1/x$ decaying function. The main difference between panels A and B is that the experimental tie-lines depend more strongly on caseinate concentrations.

3.3. Kinetics of Phase Separation. The phase separation process can usually be described in terms of two mechanisms: (a) nucleation and growth, for mixtures with concentrations in the metastable region of the phase diagram, between the binodal and spinodal and (b) spinodal decomposition, for mixtures in the unstable region above the spinodal curve. Close to the critical point, where the binodal and spinodal meet, an increase in polymer concentrations directly leads to a transition from the stable region to the spinodal region, without crossing the metastable region. To visualize spinodal decomposition in pectin/caseinate mixtures, samples close to the critical point were investigated with equal volumes of the two phases but also with 75% pectin phase/25% caseinate phase and 25% pectin phase/75% caseinate phase along the same tie-line.

As can be seen in Figure 6, after mixing was stopped structure formation was observed for the 50%/50% sample, which is characteristic of spinodal decomposition. From the interconnected structure seen in the microscopic images (left side of the insets in Figure 6A), we see the formation of the typical interconnected structure, and the Fourier transform of the images shows the appearance of a ring structure, which is typical for spinodally decomposing samples.^{26,27} The peak intensity is found

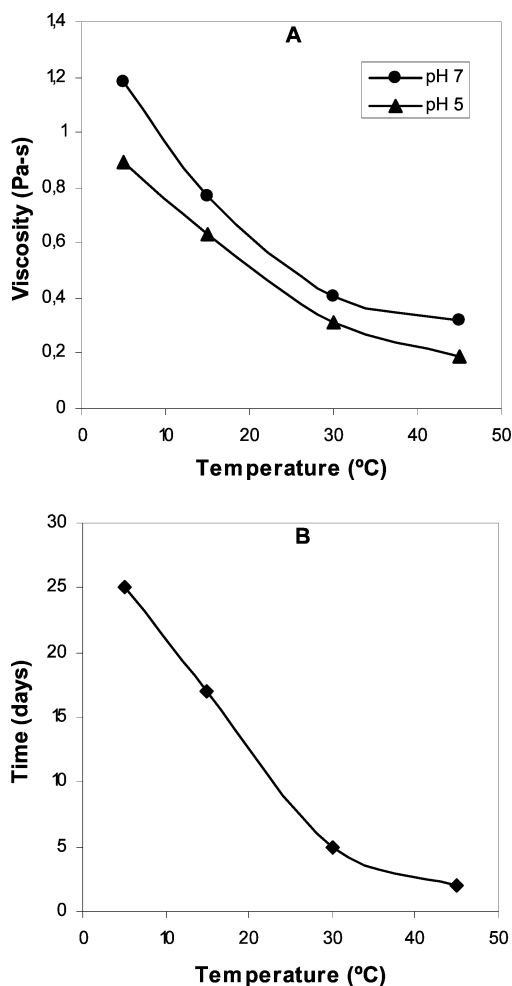


Figure 8. Viscosity of a pectin/caseinate mixture (30/30 g/L) pH 7 (panel A) and incubation time required for macroscopic phase separation (panel B) as a function of temperature.

at the typical wavelength, which we denoted as Q_{\max} . In Figure 6B, the characteristic length scale $1/Q_{\max}$ is plotted as a function of time after cessation of mixing. Shortly after mixing $1/Q_{\max}$ is apparently nearly constant, which is sometimes observed in relatively viscous demixing polysaccharide/protein mixtures.²⁷ Above $t \approx 1$ s, $1/Q_{\max}$ starts to increase significantly with time. In the regime around $t \approx 2$ s, Q_{\max} scales as $1/Q_{\max} \sim t^{0.52}$. For $t \approx 10$ s, Q_{\max} scales as $1/Q_{\max} \sim t^{0.76}$. These scaling exponentials are at variance with theoretical estimates for spinodal decomposition of $1/Q_{\max} \sim t^{1/3}$ and $1/Q_{\max} \sim t^1$ for concentrated mixtures where hydrodynamics dominates, which are both found experimentally.^{28,29}

For the 75%/25% and 25%/75% samples, even after high shear rates, the droplets did not disappear. According to Onuki,³⁰ domains that are more viscous than the continuous phase tend to be closed-shaped because the continuous phase is more easily deformed into extended shapes. Therefore, for those samples, the spinodal decomposition in the very early stages could not be visualized.

3.4. Stability. Pectin and caseinate concentrations strongly affect the phase separation rate of the biopolymer mixture, because the concentration of biopolymers influences both the thermodynamic forces in the mixture and the viscosity of the system. High polymer concentrations increase the thermodynamic incompatibility that induces the phase separation, but, on the other hand, increasing the viscosity decelerates the demixing process. Thus, the combination of both factors

determines the demixing rate. If the biopolymer concentration increases and the phase separation rate slows down, then this is due to the fact that viscous forces dominate, while thermodynamic forces dominate if the phase separation rate goes up.

Figure 7 indicates the time required for macroscopic phase separation of some pectin/caseinate mixtures at four temperatures: 5, 15, 30, and 45 °C. The samples were checked every 24 h for 1 week. It is clear that phase separation is retarded when the polymer concentration in the mixture increases (compare mixtures 20/20, 30/30, and 40/40 g/L). Indeed, when the caseinate concentration is kept constant and the pectin concentration increases, the demixing process is also retarded (see caseinate 10 and 40 g/L). In these cases the increasing viscosity of the mixtures is the dominating factor determining the demixing rate. However, when the pectin concentration is kept constant and the caseinate concentration is increased, the demixing process either slows down (see pectin 20 g/L at 15 °C) or accelerates (pectin 30 g/L at 30 and 45 °C). In the last case the thermodynamic forces dominate with respect to the rheological properties of the system.

In the late stage of spinodal decomposition, phase separation is governed by diffusion of droplets. Because of the high viscosity of concentrated polymer solutions, diffusivity of the droplets in the system decreases with increasing concentration. The low diffusivity consequently reduces the collisions between the droplets. Hence, coalescence slows down, and the demixing process may even become arrested. The effect of gelation on coarsening prevention has been reported in systems containing gelatin^{31–34} and hydroxypropyl methylcellulose.³⁵

Besides biopolymer concentration, another factor that influences the viscosity, and consequently the demixing rate, is temperature. At higher temperatures the demixing process for pectin/caseinate mixtures occurs faster. At 45 and 30 °C the macroscopic phase separation took place within 6 days. At 15 and 5 °C most of the samples were still stable after 1 week.

Figure 8A shows the viscosity of a pectin/caseinate mixture as a function of temperature. A mixture with a composition of 30/30 g/L was chosen for this experiment. The viscosities plotted were viscosities measured at a shear rate of 0.1 s^{-1} . As expected, the viscosity of the mixture increases with decreasing temperature. This change in viscosity is strongly correlated with the demixing duration (Figure 8B). Although viscosity is not the only factor that dictates the phase separation rate, it is clear that it does play an important role. Another factor that in general affects the phase separation rate is the interfacial tension,²⁸ which probably hardly changes with temperature. Depending on the storage temperature, these water-in-water mixtures can be stable for approximately a month.

3.5. Effect of pH on Pectin/Caseinate Mixtures. An interesting finding was made when the pH of the mixture was lowered to 5. Although the viscosity of the mixture was lower at that pH than at pH 7 (Figure 8A), the macroscopic phase separation took much longer (more than a week) at 45 °C. This suggests that the driving force for a segregative phase separation becomes weaker at lower pH values and thermodynamic incompatibility is suppressed. If complexation occurs, then two phases would be seen only after sedimentation of the complexes, and this could take a long time depending on the density and sizes of the complex particles.

To investigate what happens at lower pH's, the concentration of the biopolymers in both phases was measured for a 30/30 g/L mixture at different pH values. First, a mixture was prepared with pH 8, and then, by adding citric acid, the pH was brought down until the desired pH was reached. Figure 9A shows the

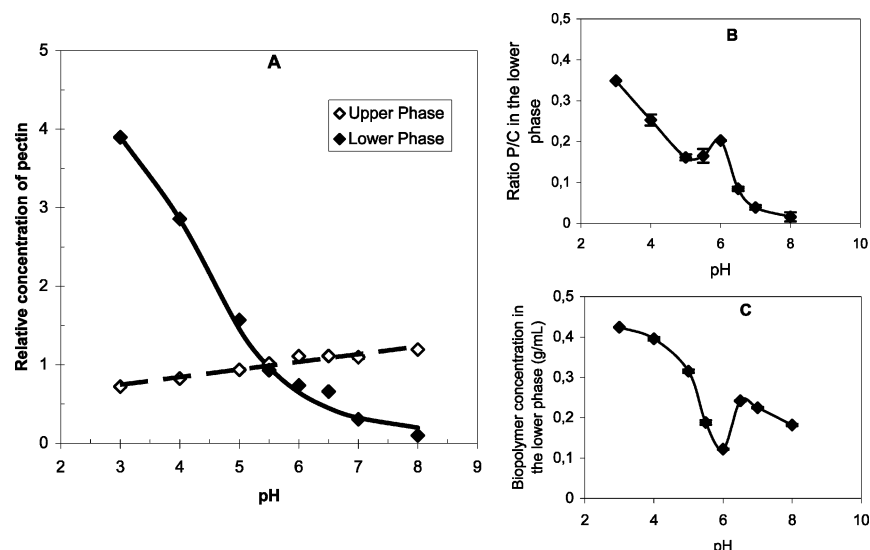


Figure 9. (A) Pectin concentration in the lower and upper phases relative to initial pectin concentration in the mixture as a function of pH. (B and C) Lower phase content at different pH values. Panel B refers to the mass ratio of pectin/caseinate, and panel C refers to the total biopolymer (pectin + caseinate) concentration in the lower phase.

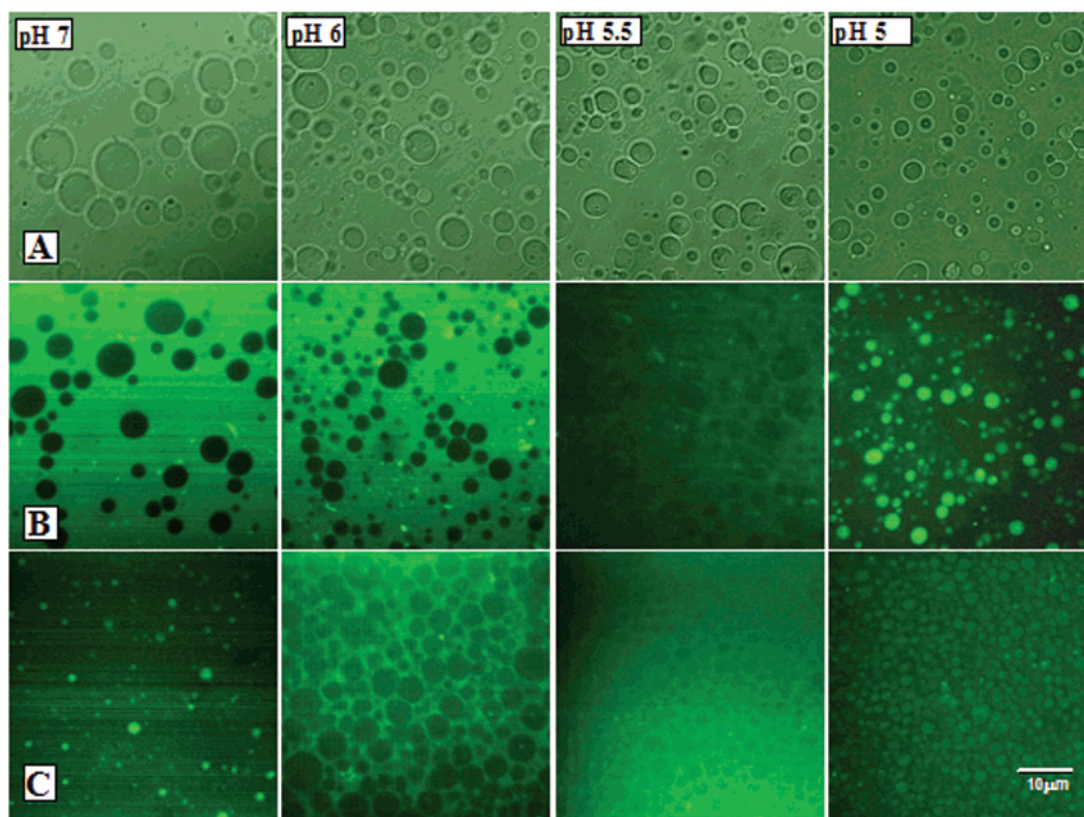


Figure 10. Transmission (row A) and CSLM images (rows B and C) of pectin/ caseinate mixtures (30/30 g/L) at different pH values. Rows A and B refer to images of the mixtures just after stirring. Row C refers to images of the lower phases after centrifugation. The scale bar for all images corresponds to 10 μm . Bright areas correspond to pectin-rich phases.

relative concentration of pectin in both the upper and the lower phases compared to the initial concentration as a function of pH.

At high pH values the concentration of pectin in the upper phase is higher than the concentration in the original mixture. In the lower phase, only small amounts of the polysaccharide are found. This reflects a phase separation driven by segregative interactions. As the pH is lowered, the concentration of pectin in the upper phase decreases while the concentration increases in the lower phase. Because the lower phase volume is much smaller than the upper phase volume (it can be 10 times smaller

depending on the pH), the concentration in the lower phase increases strongly with decreasing pH. The crossover point at pH 5.5 indicates that at this pH the concentration of pectin is equal in both phases and that below this point pectin becomes more concentrated in the lower phase. In contrary to pectin, caseinate remained more concentrated in the lower phase independent of the pH. These results suggest that pectin and caseinate concentrate in a single phase below pH 5.5, which is characteristic for complexation.

The reason why at pH > 6 pectin/caseinate mixtures phase-separate is that the polymers repel each other. Under those

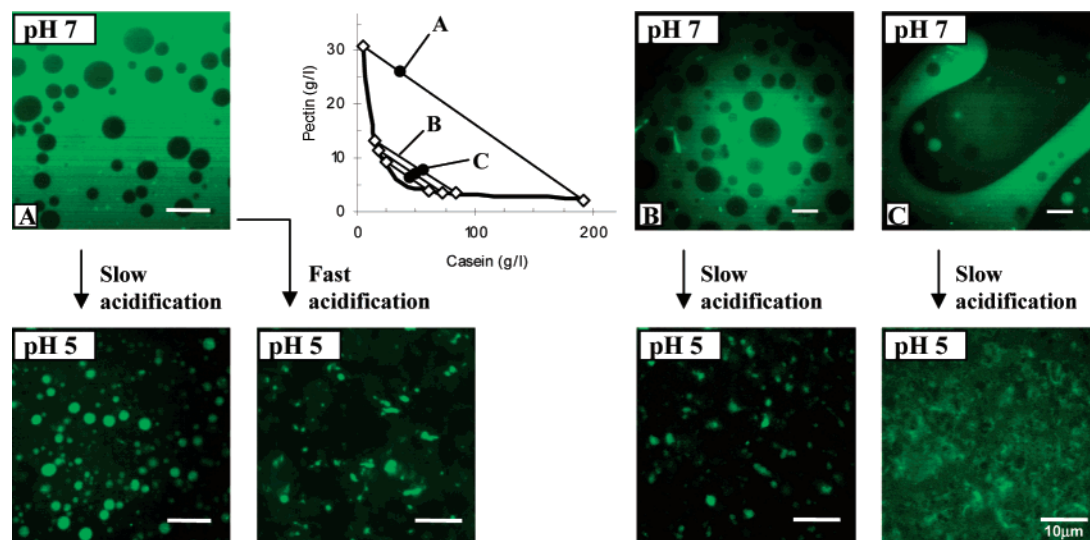


Figure 11. CLSM images of pectin/caseinate mixtures at three different points in the phase diagram, before and after acidification. The scale bars correspond to 10 μm . Bright areas correspond to pectin-rich phases.

conditions both pectin, with pK_a 's of carboxylic groups close to 3, and caseinate, with $\text{pI} \approx 4.6$, carry negative charges, which makes repulsive electrostatic and excluded volume interactions dominant. In that case, incompatibility takes place driven by depletion forces. However, as the pH of the mixture is lowered, caseinate becomes less negative. Even though the protein carries a net negative charge above the pI , the presence of regions on the caseinate molecules that are positively charged even at high pH's can attract the negatively charged pectins. This attraction becomes stronger with decreasing pH below the isoelectric point, when the amphoteric caseinate becomes positively charged.

Figure 10 shows images of the 30/30 g/L mixture in transmission and confocal microscopy. At pH's 7 and 6 the mixtures (rows A and B) consisted of caseinate-rich droplets in a pectin-rich continuous phase. At pH 5.5 confocal contrast is negligible, which means that the concentration of pectin in the droplets and in the continuous phase becomes similar. This agrees with the results of Figure 9A where the upper phase concentration curve meets the lower phase close to pH 5.5. Further lowering of the pH (pH 5) pushes more pectin into the droplets, and they become brighter than the outside medium. The "migration" of pectin from the continuous phase to the particles explains the lower viscosity of the mixture at pH 5 compared with pH 7 (Figure 8A).

Images of the lower phases after phase separation (Figure 10C) were also investigated. At pH 7, the lower phase was mainly caseinate-enriched with some small pectin-rich droplets. At pH 6, the lower phase was not a continuous phase but rather caseinate-rich particles interconnected by pectin-rich areas, indicating that coalescence was prevented. At pH 5.5 contrast nearly vanishes, while particles become distinguishable again at pH 5. Figures 9B and 9C show the lower phase content at different pH values. In panel B, we see that the mass ratio of pectin/caseinate increases with decreasing pH. At pH 6, although pectin is not concentrated in the particles, the amount in the lower phase increases because it stays entrapped between the particles (Figure 10C). This causes an increase in the volume of the lower phase, which explains the reduction of total biopolymer concentration in the lower phase at that pH (Figure 9C). Visual observations of the mixtures also indicate that complexation occurs at pH values smaller than 6. Between pH 8 and 6.5, the mixture was transparent and so were the two phases after demixing. Starting from pH 6, the mixture became

more turbid, and at pH 5 it was opalescent white, indicating the presence of insoluble matter (complexes). At this point, the bottom phase after centrifugation was solid and not liquid as for high pH mixtures.

The slow acidification of polysaccharide/protein mixtures from a pH where no associative interaction occurs has been performed before by using glucono- δ -lactone (GDL) to study electrostatic adsorption of pectins onto casein micelles in milk¹⁰ and electrostatic complexation between β -lactoglobulin and xanthan,³⁶ β -lactoglobulin and pectins,³⁷ and β -lactoglobulin and Acacia gum.³⁸ However, the use of xanthan or pectins as polyanionic biopolymers did not induce the formation of coacervates but resulted in rather small¹⁰ or fractal aggregates with fractal dimensions in the range of 1.7–2.5.³⁷ Liquid droplets or vesicles were observed only with acacia gum.^{39,40}

In our case, spherical particles $\sim 3 \mu\text{m}$ in diameter were obtained after lowering the pH slowly (~ 1 drop of citric acid/min) from 7 to 5 for mixtures with high pectin concentrations (Figure 11, situation A). When the pH was lowered fast (a single drop of citric acid/s), the particles lost sphericity. It seems that pectin needs time to diffuse into the droplets and adsorb onto micellar caseinates and stabilize them. When the acidification is performed rapidly, not all micelles become stabilized by pectin molecules, and some of them aggregate, resulting in nonspherical particles. This also happens when the concentration of pectin in the mixture is reduced (Figure 11, situations B and C). In such cases, the amount of pectin is insufficient to stabilize all caseinate micelles, and even with slow acidification, the particles become nonspherical. For samples with volume fractions of 50% (situation C in Figure 11) and 25% pectin/75% caseinate (not shown), the concentration of pectin was, compared to caseinate concentration, low to such a degree that, at pH 5, pectin was still concentrated in the continuous phase (the particles were dark). Because the amount of pectin available to complex with caseinate was so small, caseinates needed to have more positively charged groups (pH below the isoelectric point) to attract more pectin, so that the particles would become more concentrated in pectin than the continuous phase was. It is important to mention that this complexation in pectin/caseinate mixtures is reversible. As the pH is increased again, the pectin molecules migrate from the particles to the bulk phase, and the droplets become flexible and start coalescing as happens at high pH values.

Despite the pH sensitivity, pectin/caseinate complexes are found to be insensitive to salt addition. This contrasts with the behavior of other studies on polysaccharide/protein complexes.^{41,42} Even at high ionic strengths (100 mM NaCl) the particles did not change into soluble complexes or to a segregation of pectin and caseinate. It is a well-known phenomenon that the microions present in the solution screen the charges of the polymers and thus reduce the range of their associative interactions.² In pectin/caseinate complexes, maybe because of the high biopolymer concentration used (total 60 g/L), the amount of microions was below that required for complexation suppression.

Conclusions

In this paper we have shown that pH and biopolymer concentration play an important role in the phase stability of pectin/caseinate mixtures. A phase diagram constructed at pH 7, where both polymers are negatively charged, reveals that the mixture is stable only at small biopolymer concentrations. Higher concentrations lead to phase separation, due to thermodynamic incompatibility, and formation of a water-in-water emulsion. We could confirm that thermodynamic incompatibility occurs at pH > 6, because the phase line in that regime is insensitive to temperature and the observed phase diagram could be rationalized by comparing with theory for a dispersion of hard spheres (mimicking micellar caseinates) plus polymer chains (mimicking pectin polymers) in solution with excluded volume interactions between the segments. Spinodal decomposition was observed in the early stages of the phase separation. The phase separation rate slows down as the concentration of pectin in the system is increased, due to the higher viscosity of those mixtures. Because viscosity changes with temperature, the stability of the emulsions is also dependent on the storage temperature.

When the pH of the incompatible mixture is lowered below 6, complexation takes place, because caseinate then becomes significantly positively charged and attracts the negatively charged pectin molecules, resulting in the formation of large pectin/caseinate particles that are resistant to high ionic strengths. These particles are more spherical at high pectin concentrations and when the acidification is performed slowly, because of better stabilization of micellar caseinates by pectin at these conditions. The complexation process is reversible so that the particles turn into soluble droplets (emulsion) again with increasing pH. This versatility of pectin/caseinate mixtures could be relevant especially for pharmaceutical applications in the field of microencapsulation of active compounds and drug delivery systems. In the particular case of oral administration, where the pH changes from acidic in the stomach to neutral at the intestine, pectin/caseinate particles would be excellent candidates.

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